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energy is defined as  
the amount of energy

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required to break a bond. These values are positive, indicating that bond breaking is endothermic. Bond energies reported in kilojoules per mole (kJ/mol). The energy for breaking a hydrogen-hydrogen bond is 436

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example  $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \rightarrow 2\text{HF}(\text{g})$ .  
 $\Delta H_{\text{rxn}} = 2$   
[energy used for  
breaking bonds] -  
[energy formed in  
making bonds] [436  
kJ/mol + 155 kJ/mol] -  
[2(567 kJ/mol)] = - 543  
kJ/mol. Bond energy is  
defined as the amount  
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break a bond.

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Bond energy: 436

kJ/mol Bond Energy:

155 kJ/mol Bond

Energy: 567 kJ/mol

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H= Bonds Broken

Bonds Formed = [H-H  
+ Cl-Cl]-[2(H-Cl)] =

[436

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The average bond energies in Table T3 are the averages of bond dissociation energies. For example the average bond energy of O-H in  $\text{H}_2\text{O}$  is 464 kJ/mol. This is due to the fact that the H-OH bond requires 498.7 kJ/mol to dissociate, while the O-H bond needs 428 kJ/mol.

## 3.7: Bond Energies &

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Bond Energy Explained

. Bond energy is based on an average of bond dissociation values for species in the gas phase, typically at a temperature of 298 Kelvin. It may be found by measuring or calculating the enthalpy change of breaking a molecule into its component atoms and ions and

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dividing the value by  
the number of  
chemical bonds.

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